Q1.

Which row correctly shows general trends in properties across Period 3?

	Atomic radius	First ionisation energy	
A	Decreases	Increases	⬭
B	Decreases	Decreases	⬭
C	Increases	Increases	⬭
D	Increases	Decreases	⬭

(Total 1 mark)

Q2.

Consider the Period 3 elements from sodium to chlorine.

Which statement is correct?

A Sodium has the smallest atomic radius. ⬭

B Aluminium has the highest melting point. ⬭

C Sulfur is the most electronegative. ⬭

D Chlorine has the highest first ionisation energy. ⬭

(Total 1 mark)

Q3.

Which statement is **not** correct about the Period 3 elements sodium to chlorine?

A Sodium has the largest atomic radius. ⬭

B Sodium has the lowest melting point. ⬭

C Silicon has the highest melting point. ⬭

D Chlorine has the highest first ionisation energy. ⬭

(Total 1 mark)

Q4.

Which block in the Periodic Table contains the element samarium (Sm)?

A d block ⬭

B f block ⬭

C p block ⬭

D s block ⬭

(Total 1 mark)

Q5.

Which statement is correct about the Group 1 elements?

A The Cs^+ ion has a more negative enthalpy of hydration than the Rb^+ ion. ⬭

B The enthalpy of atomisation for potassium is greater than the enthalpy of atomisation for sodium. ⬭

C The melting point of potassium is higher than the melting point of sodium. ⬭

D The second ionisation energy of rubidium is lower than the second ionisation energy of lithium. ⬭

(Total 1 mark)